

Acta Crystallographica Section E

Structure Reports

Online

ISSN 1600-5368

4-Chloro-2-[(*E*)-(4-nitrophenyl)diazenyl]-phenol

Leonid A. Aslanov,* Ksenia A. Paseshnichenko and Alexandr V. Yatsenko

Department of Chemistry, Moscow State University, 119992 Moscow, Russian Federation

Correspondence e-mail: aslanov@struct.chem.msu.ru

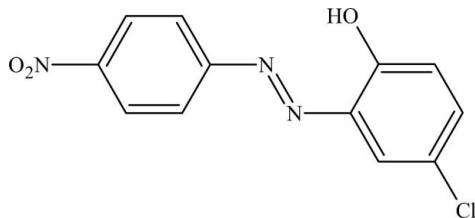
Received 14 January 2009; accepted 29 January 2009

Key indicators: single-crystal X-ray study; $T = 291$ K; mean $\sigma(\text{C}-\text{C}) = 0.002$ Å; R factor = 0.037; wR factor = 0.098; data-to-parameter ratio = 14.8.

The title compound, $\text{C}_{12}\text{H}_8\text{ClN}_3\text{O}_3$, in the crystalline state and in solution, exists in the azo form, as predicted by density functional theory (DFT) calculations. The molecule is approximately planar [the dihedral angle between the rings is 1.83 (8°)], with the nitro group slightly twisted [13.4 (2°)] relative to the benzene ring. Translationally related molecules form stacks along [010] with an interplanar distance of 3.400 (2) Å. The hydroxy group forms an intramolecular hydrogen bond with the azo N atom.

Related literature

For the crystal structure of a closely related molecule, (1*Z*)-4-hydroxybenzo-1,2-quinone-1-[(2-chloro-4-nitrophenyl)-hydrazone], that crystallizes as a hydrazone tautomer, see: You *et al.* (2004). For reference structural data, see: Allen (2002). For details of the synthetic procedure, see: Fierz-David & Blangey (1949). For background on DFT calculations, see: Becke (1993); Klamt & Schüürmann (1993); Krishnan *et al.* (1980); Lee *et al.* (1988); Schmidt *et al.* (1993). For the concept of resonance-assisted hydrogen bonds, see: Gilli *et al.* (1989).



Experimental

Crystal data

 $\text{C}_{12}\text{H}_8\text{ClN}_3\text{O}_3$
 $M_r = 277.66$

 Monoclinic, $P2_1/c$
 $a = 19.008$ (5) Å

 $b = 4.817$ (2) Å
 $c = 12.862$ (4) Å
 $\beta = 92.65$ (2°)
 $V = 1176.4$ (7) Å³
 $Z = 4$

 Mo $K\alpha$ radiation
 $\mu = 0.33$ mm⁻¹
 $T = 291$ (2) K
 $0.40 \times 0.20 \times 0.15$ mm

Data collection

 Enraf-Nonius CAD-4
 diffractometer
 Absorption correction: none
 2567 measured reflections
 2567 independent reflections

 2110 reflections with $I > 2\sigma(I)$
 3 standard reflections
 frequency: 90 min
 intensity decay: 4%

Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.037$
 $wR(F^2) = 0.098$
 $S = 1.56$
 2567 reflections

 173 parameters
 H-atom parameters constrained
 $\Delta\rho_{\text{max}} = 0.25$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.14$ e Å⁻³
Table 1
 Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
O1-H1 \cdots N2	0.82	1.88	2.5777 (17)	143

Data collection: *CAD-4 Software* (Enraf-Nonius, 1989); cell refinement: *CAD-4 Software*; data reduction: *PROFIT* (Streltsov & Zavodnik, 1989) routine of *WinGX* (Farrugia, 1999); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *PLATON* (Spek, 2003); software used to prepare material for publication: *PLATON*.

KAP and AVY thank the ICDD for financial assistance (grant No. 93-05).

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: GK2186).

References

- Allen, F. H. (2002). *Acta Cryst.* **B58**, 380–388.
 Becke, A. D. (1993). *J. Chem. Phys.* **98**, 5648–5652.
 Enraf-Nonius (1989). *CAD-4 Software*. Enraf-Nonius, Delft, The Netherlands.
 Farrugia, L. J. (1999). *J. Appl. Cryst.* **32**, 837–838.
 Fierz-David, H. E. & Blangey, L. (1949). *Fundamental Processes of Dye Chemistry*, pp. 236–240. London: Interscience.
 Gilli, G., Belucci, F., Ferretti, V. & Bertolasi, V. (1989). *J. Am. Chem. Soc.* **111**, 1023–1028.
 Klamt, A. & Schüürmann, G. (1993). *J. Chem. Soc. Perkin Trans. 2*, pp. 799–805.
 Krishnan, R., Binkley, J. S., Seeger, R. & Pople, J. A. (1980). *J. Chem. Phys.* **72**, 650–654.
 Lee, C., Yang, W. & Parr, R. G. (1988). *Phys. Rev.* **B37**, 785–789.
 Schmidt, M. W., Baldrige, K. K., Boatz, J. A., Elbert, S. T., Gordon, M. S., Jensen, J. J., Koseki, S., Matsunaga, N., Nguyen, K. A., Su, S., Windus, T. L., Dupuis, M. & Montgomery, J. A. (1993). *J. Comput. Chem.* **14**, 1347–1363.
 Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
 Spek, A. L. (2003). *J. Appl. Cryst.* **36**, 7–13.
 Streltsov, V. A. & Zavodnik, V. E. (1989). *Sov. Phys. Crystallogr.* **34**, 824–828.
 You, X.-L., Lu, C.-R., Huang, Z.-L. & Zhang, D.-C. (2004). *Dyes Pigm.* **63**, 217–223.

supplementary materials

Acta Cryst. (2009). E65, o497 [doi:10.1107/S1600536809003675]

4-Chloro-2-[(*E*)-(4-nitrophenyl)diazenyl]phenol

L. A. Aslanov, K. A. Paseshnichenko and A. V. Yatsenko

Comment

The present work was fulfilled in the course of study of the hydroxyazo-ketohydrazone tautomerism in phenylazophenols. The title compound, (I), exists in crystals as the azo form (Fig. 1). It is evidenced, firstly, by the fact that the H atom was found and refined in the vicinity of O atom, and secondly, by comparison of the molecular geometry with numerous structures of azo tautomers found in the Cambridge Structural Database (Allen, 2002). Since the UV-visible spectra of the crystalline title compound resemble its spectra in solutions, the azo tautomer has to predominate in solutions as well.

However, recently it has been reported that (1*Z*)-4-hydroxybenzo-1,2-quinone 1-[(2-chloro-4-nitrophenyl)hydrazone (II), the compound closely related to (I), exists in crystals as the hydrazone tautomer (You *et al.*, 2004).

The azo-hydrazone equilibrium is known to be shifted by the effect of donor and acceptor substituents and also by intra- and intermolecular hydrogen bonds. In order to evaluate the relative importance of these factors, we have performed the DFT calculations of azo and hydrazone tautomers of (I) and (II). Calculations were carried out using GAMESS (Schmidt *et al.*, 1993) with B3LYP exchange-correlation functional (Becke, 1993; Lee *et al.*, 1988) and 6-311G** basis set (Krishnan *et al.*, 1980). After geometry of an isolated molecule has been optimized, molecular structure was fixed, and the effect of nonspecific intermolecular interactions was accounted by COSMO method (Klamt & Schüürmann, 1993), taking the dielectric permeability equal to 10. The results indicate that for 2-phenyldiazenylphenol (III), the azo form is by 10.5 kJ/mol more stable than the hydrazone form. For compound (I), this difference decreases to 7.5 kJ/mol and for (II) - to 6.8 kJ/mol, but nonetheless the azo form is still preferable.

Thus, the difference between (I) and (II) most probably arises from specific intermolecular interactions. In (I), there is the only worthnoting intermolecular contact C15—H15 \cdots O2 (-*x*, -1 - *y*, -*z*) (H15 \cdots O2 2.56 Å, C15 \cdots O2 3.452 (2) Å, C15—H15 \cdots O2 161°), which cannot have any effect on the relative stability of tautomers. In (II), the keto group forms a strong hydrogen bond with the hydroxy group of a neighboring molecule (O \cdots H 1.74 Å, O \cdots O 2.581 (2) Å, O—H \cdots O 173°). This interaction stabilizes the hydrazone tautomer, according to the conception of resonance-assisted hydrogen bonds (Gilli *et al.*, 1989). So, the shift of tautomeric equilibrium in (II) towards the hydrazone form should be most probably rationalized by formation of intermolecular hydrogen bonds.

Experimental

The title compound was prepared by coupling of *p*-nitrophenyldiazonium chloride with *p*-chlorophenol. For details of the synthetic procedure, see Fierz-David & Blangey (1949). Single crystals were grown by slow evaporation of ethanol solution.

Refinement

H atoms were located in a difference map and refined freely, but at final stage they were positioned geometrically and refined using a riding model with C—H = 0.93 Å, O—H = 0.82 Å and with $U_{\text{iso}}(\text{H}) = 1.2$ times $U_{\text{eq}}(\text{C})$, $U_{\text{iso}}(\text{H}) = 1.5$ times $U_{\text{eq}}(\text{O})$

Figures



Fig. 1. The molecular structure of the title compound with atomic labels and 50% probability displacement ellipsoids for non-H atoms.

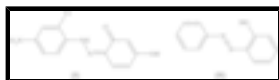


Fig. 2. Chemical diagrams of (II) and (III).

4-Chloro-2-[(E)-(4-nitrophenyl)diazenyl]phenol

Crystal data

$C_{12}H_8ClN_3O_3$

$M_r = 277.66$

Monoclinic, $P2_1/c$

Hall symbol: -P 2ybc

$a = 19.008$ (5) Å

$b = 4.817$ (2) Å

$c = 12.862$ (4) Å

$\beta = 92.65$ (2)°

$V = 1176.4$ (7) Å³

$Z = 4$

$F_{000} = 568$

$D_x = 1.568$ Mg m⁻³

Mo $K\alpha$ radiation

$\lambda = 0.71073$ Å

Cell parameters from 25 reflections

$\theta = 16.8$ – 18.8 °

$\mu = 0.33$ mm⁻¹

$T = 291$ K

Prism, red

$0.40 \times 0.20 \times 0.15$ mm

Data collection

Enraf–Nonius CAD-4
diffractometer

Radiation source: fine-focus sealed tube

Monochromator: graphite

$T = 291$ K

nonprofiled ω scans

Absorption correction: none

2567 measured reflections

2567 independent reflections

2110 reflections with $I > 2\sigma(I)$

$R_{int} = 0.0$

$\theta_{max} = 27.0$ °

$\theta_{min} = 1.1$ °

$h = -24 \rightarrow 24$

$k = 0 \rightarrow 6$

$l = 0 \rightarrow 16$

3 standard reflections

every 90 min

intensity decay: 4%

Refinement

Refinement on F^2

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.037$

$wR(F^2) = 0.098$

$S = 1.56$

Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring sites

H-atom parameters constrained

$w = 1/[\sigma^2(F_o^2) + (0.04P)^2]$

where $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{max} = 0.001$

2567 reflections $\Delta\rho_{\max} = 0.25 \text{ e } \text{\AA}^{-3}$
 173 parameters $\Delta\rho_{\min} = -0.14 \text{ e } \text{\AA}^{-3}$
 Primary atom site location: structure-invariant direct methods Extinction correction: none

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
C11	0.463076 (19)	1.28010 (9)	0.16237 (3)	0.06275 (16)
O1	0.27449 (6)	0.7166 (3)	-0.14202 (8)	0.0703 (4)
H1	0.2491	0.6045	-0.1138	0.105*
O2	-0.00077 (6)	-0.4250 (3)	0.11633 (9)	0.0665 (3)
O3	0.05796 (6)	-0.4261 (2)	0.26393 (8)	0.0592 (3)
N1	0.27334 (5)	0.5696 (2)	0.07974 (9)	0.0436 (3)
N2	0.23136 (6)	0.4462 (2)	0.01583 (9)	0.0450 (3)
N3	0.04760 (6)	-0.3453 (2)	0.17452 (10)	0.0468 (3)
C1	0.31710 (7)	0.8384 (3)	-0.07021 (11)	0.0496 (4)
C2	0.31686 (6)	0.7693 (3)	0.03656 (10)	0.0420 (3)
C3	0.36225 (7)	0.9076 (3)	0.10740 (11)	0.0444 (3)
H3	0.3620	0.8642	0.1778	0.053*
C4	0.40705 (7)	1.1065 (3)	0.07365 (12)	0.0475 (3)
C5	0.40778 (8)	1.1751 (3)	-0.03104 (12)	0.0556 (4)
H5	0.4387	1.3097	-0.0534	0.067*
C6	0.36281 (8)	1.0440 (4)	-0.10150 (12)	0.0589 (4)
H6	0.3629	1.0935	-0.1714	0.071*
C11	0.18678 (6)	0.2481 (3)	0.06134 (10)	0.0402 (3)
C12	0.18713 (7)	0.1938 (3)	0.16816 (11)	0.0472 (3)
H12	0.2177	0.2884	0.2143	0.057*
C13	0.14144 (7)	-0.0020 (3)	0.20368 (10)	0.0473 (3)
H13	0.1412	-0.0428	0.2743	0.057*
C14	0.09613 (6)	-0.1372 (3)	0.13447 (10)	0.0401 (3)
C15	0.09456 (7)	-0.0863 (3)	0.02927 (10)	0.0456 (3)
H15	0.0634	-0.1801	-0.0161	0.055*
C16	0.14084 (7)	0.1087 (3)	-0.00691 (11)	0.0469 (3)
H16	0.1411	0.1465	-0.0778	0.056*

supplementary materials

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C11	0.0544 (2)	0.0552 (3)	0.0784 (3)	-0.00993 (18)	0.00095 (19)	-0.0052 (2)
O1	0.0844 (8)	0.0812 (9)	0.0452 (6)	-0.0262 (7)	0.0024 (5)	0.0036 (6)
O2	0.0716 (7)	0.0648 (8)	0.0636 (7)	-0.0297 (6)	0.0068 (6)	-0.0067 (6)
O3	0.0716 (7)	0.0500 (6)	0.0570 (7)	-0.0031 (5)	0.0129 (5)	0.0118 (5)
N1	0.0437 (6)	0.0373 (6)	0.0502 (6)	0.0005 (5)	0.0071 (5)	0.0018 (5)
N2	0.0460 (6)	0.0406 (6)	0.0486 (6)	-0.0002 (5)	0.0054 (5)	0.0021 (5)
N3	0.0548 (7)	0.0348 (6)	0.0518 (7)	0.0001 (5)	0.0128 (5)	-0.0031 (5)
C1	0.0523 (8)	0.0494 (8)	0.0477 (8)	0.0001 (7)	0.0095 (6)	0.0004 (7)
C2	0.0418 (6)	0.0367 (7)	0.0482 (8)	0.0030 (6)	0.0092 (6)	0.0027 (6)
C3	0.0452 (7)	0.0382 (7)	0.0502 (7)	0.0040 (6)	0.0072 (6)	0.0035 (6)
C4	0.0432 (7)	0.0373 (7)	0.0626 (9)	0.0023 (6)	0.0067 (6)	-0.0042 (7)
C5	0.0551 (8)	0.0473 (9)	0.0663 (10)	-0.0056 (7)	0.0214 (7)	0.0042 (8)
C6	0.0663 (9)	0.0596 (10)	0.0522 (9)	-0.0060 (8)	0.0176 (7)	0.0075 (8)
C11	0.0402 (6)	0.0351 (7)	0.0456 (7)	0.0033 (5)	0.0055 (5)	0.0018 (6)
C12	0.0478 (7)	0.0485 (8)	0.0451 (8)	-0.0054 (6)	0.0000 (6)	-0.0017 (6)
C13	0.0537 (8)	0.0501 (8)	0.0384 (7)	-0.0021 (7)	0.0041 (6)	0.0024 (6)
C14	0.0419 (6)	0.0296 (6)	0.0492 (7)	0.0025 (5)	0.0081 (5)	-0.0005 (6)
C15	0.0505 (7)	0.0426 (8)	0.0432 (7)	-0.0058 (6)	-0.0015 (6)	-0.0025 (6)
C16	0.0540 (8)	0.0472 (8)	0.0395 (7)	-0.0023 (6)	0.0030 (6)	0.0053 (6)

Geometric parameters (\AA , $^\circ$)

C11—C4	1.7387 (16)	C4—C5	1.387 (2)
O1—C1	1.3358 (17)	C5—C6	1.371 (2)
O1—H1	0.8200	C5—H5	0.9300
O2—N3	1.2203 (15)	C6—H6	0.9300
O3—N3	1.2214 (15)	C11—C16	1.3831 (18)
N1—N2	1.2670 (16)	C11—C12	1.3983 (19)
N1—C2	1.4001 (17)	C12—C13	1.3743 (19)
N2—C11	1.4203 (17)	C12—H12	0.9300
N3—C14	1.4712 (17)	C13—C14	1.3741 (19)
C1—C6	1.389 (2)	C13—H13	0.9300
C1—C2	1.413 (2)	C14—C15	1.3741 (19)
C2—C3	1.3949 (19)	C15—C16	1.3823 (19)
C3—C4	1.3660 (19)	C15—H15	0.9300
C3—H3	0.9300	C16—H16	0.9300
C1—O1—H1	109.5	C5—C6—C1	121.06 (14)
N2—N1—C2	115.52 (11)	C5—C6—H6	119.5
N1—N2—C11	114.69 (11)	C1—C6—H6	119.5
O2—N3—O3	124.14 (12)	C16—C11—C12	120.47 (12)
O2—N3—C14	117.79 (12)	C16—C11—N2	115.84 (12)
O3—N3—C14	118.06 (12)	C12—C11—N2	123.69 (12)
O1—C1—C6	118.70 (13)	C13—C12—C11	118.72 (13)
O1—C1—C2	122.63 (13)	C13—C12—H12	120.6

C6—C1—C2	118.67 (14)	C11—C12—H12	120.6
C3—C2—N1	115.33 (12)	C14—C13—C12	119.79 (12)
C3—C2—C1	119.50 (13)	C14—C13—H13	120.1
N1—C2—C1	125.17 (13)	C12—C13—H13	120.1
C4—C3—C2	120.21 (13)	C13—C14—C15	122.53 (12)
C4—C3—H3	119.9	C13—C14—N3	118.73 (12)
C2—C3—H3	119.9	C15—C14—N3	118.75 (12)
C3—C4—C5	120.66 (14)	C14—C15—C16	117.88 (12)
C3—C4—C11	120.05 (12)	C14—C15—H15	121.1
C5—C4—C11	119.29 (11)	C16—C15—H15	121.1
C6—C5—C4	119.88 (14)	C15—C16—C11	120.61 (12)
C6—C5—H5	120.1	C15—C16—H16	119.7
C4—C5—H5	120.1	C11—C16—H16	119.7
C2—N1—N2—C11	-178.90 (10)	N1—N2—C11—C16	-179.47 (11)
N2—N1—C2—C3	179.08 (11)	N1—N2—C11—C12	1.27 (18)
N2—N1—C2—C1	-0.30 (19)	C16—C11—C12—C13	0.5 (2)
O1—C1—C2—C3	-179.22 (13)	N2—C11—C12—C13	179.77 (12)
C6—C1—C2—C3	-0.1 (2)	C11—C12—C13—C14	-0.7 (2)
O1—C1—C2—N1	0.1 (2)	C12—C13—C14—C15	0.4 (2)
C6—C1—C2—N1	179.24 (13)	C12—C13—C14—N3	-179.82 (12)
N1—C2—C3—C4	179.97 (11)	O2—N3—C14—C13	166.51 (12)
C1—C2—C3—C4	-0.6 (2)	O3—N3—C14—C13	-13.07 (17)
C2—C3—C4—C5	0.4 (2)	O2—N3—C14—C15	-13.72 (18)
C2—C3—C4—C11	179.53 (10)	O3—N3—C14—C15	166.71 (12)
C3—C4—C5—C6	0.5 (2)	C13—C14—C15—C16	0.1 (2)
C11—C4—C5—C6	-178.60 (12)	N3—C14—C15—C16	-179.63 (11)
C4—C5—C6—C1	-1.3 (2)	C14—C15—C16—C11	-0.3 (2)
O1—C1—C6—C5	-179.81 (14)	C12—C11—C16—C15	0.0 (2)
C2—C1—C6—C5	1.1 (2)	N2—C11—C16—C15	-179.28 (12)

Hydrogen-bond geometry (Å, °)

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
O1—H1...N2	0.82	1.88	2.5777 (17)	143

Fig. 1

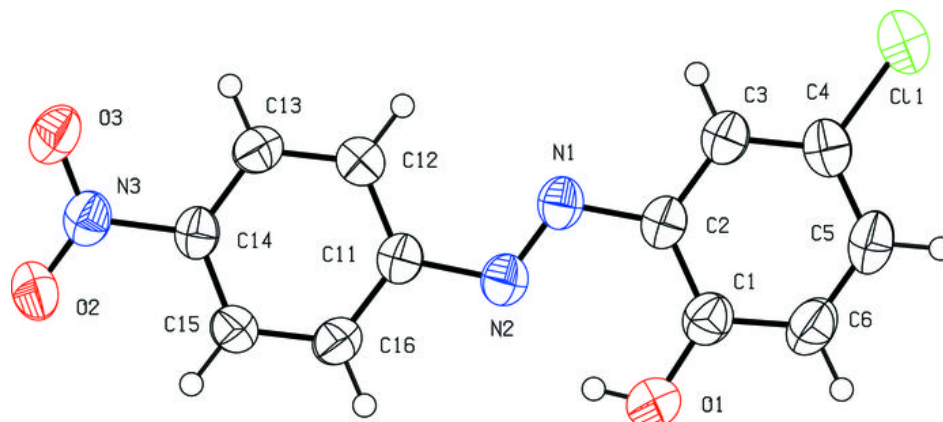


Fig. 2

